To Balance a Redox Reaction:

1. Determine the electron donor and electron acceptor participating in the reaction by assessing the oxidation states of each element.

2. Write an oxidation half-reaction for the electron donor and a reduction half-reaction for the electron acceptor: In each half reaction:
   a. Balance the redox elements.
   b. Determine the number of electrons transferred.
   c. Balance all other elements besides O and H.
   d. Balance O by adding H₂O to the other side.
   e. Balance H by adding H⁺ to other side.
   f. Check for charge-balance in the half reaction. If not, check charged species for problem.

3. Multiply the two half-reactions so that the electrons produced in the oxidation reaction equal the electrons consumed in the reduction reaction.

4. Add half-reactions to get the overall reaction. The electrons should DROP OUT of the reaction. Eliminate species that appear on both sides of the reaction by subtraction.